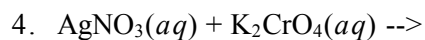
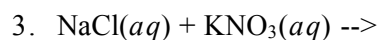
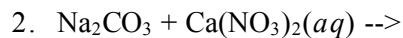
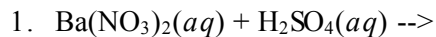


**Stoichiometry Worksheet**

Identify the solid product that forms when the following aqueous solutions are mixed:



An aqueous solution of ammonium sulfate is allowed to react with an aqueous solution of lead(II) nitrate.

- \_\_\_\_\_ 5. Identify the solid in the balanced equation.
- |                               |   |
|-------------------------------|---|
| A. $\text{Pb}(\text{NO}_3)_2$ | C. There is no solid formed when the two solutions are mixed. |
| B. $\text{PbSO}_4$            | D. $(\text{NH}_4)_2\text{SO}_4$                               |
- \_\_\_\_\_ 6. What is the coefficient of the solid in the balanced equation (in standard form)?
- |      |      |
|------|------|
| A. 3 | C. 1 |
| B. 4 | D. 6 |
- \_\_\_\_\_ 7. The complete ionic equation contains which of the following species (when balanced in standard form)?
- |                            |                          |
|----------------------------|--------------------------|
| A. $2\text{SO}_4^{2-}(aq)$ | C. $\text{NO}_3^-(aq)$   |
| B. $2\text{NH}_4^+(aq)$    | D. $3\text{Pb}^{2+}(aq)$ |
- \_\_\_\_\_ 8. The net ionic equation contains which of the following species (when balanced in standard form)?
- |                         |                            |
|-------------------------|----------------------------|
| A. $\text{NO}_3^-(aq)$  | C. $2\text{SO}_4^{2-}(aq)$ |
| B. $\text{Pb}^{2+}(aq)$ | D. $2\text{NO}_3^-(aq)$    |